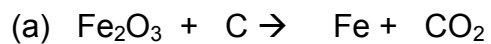


Working with Chemical Equations

CHM 1311.003 - Practice Questions

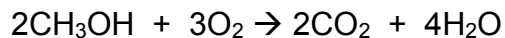
1. Balance:



2. Given the equation below, how many kg of O_2 can be produced from 1.0 kg of KClO_3 ?

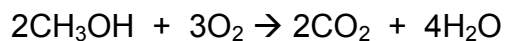


3. Combustion of methyl alcohol in oxygen according to the equation:



was conducted. When 6.40 g of CH_3OH and 10.2 g of O_2 were mixed and ignited, 6.12 g of CO_2 was obtained. What was the percent yield of CO_2 ?

4. Combustion of methyl alcohol in oxygen according to the equation:



was conducted. When 24.0 g of CH_3OH and 48.0 g of O_2 were mixed and ignited, 18.2 g of CO_2 was obtained. Based on methyl alcohol, what was the percent yield of CO_2 ?